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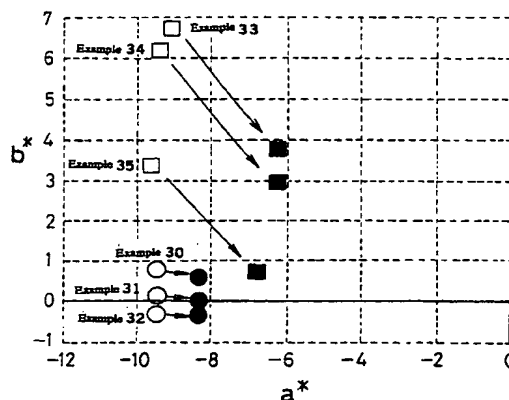
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(54) Ultraviolet/infrared absorbent low transmittance glass

(57) Ultraviolet/infrared absorbent low transmittance glass exhibiting having an almost neutral color such as turquoise blue and deep green and exhibiting low or middle visible light transmittance, low total solar energy transmittance, and low ultraviolet transmittance consists of base glass comprising: 65 to 80 wt. % SiO₂; 0 to 5 wt. % Al₂O₃; 0 to 10 wt. % MgO; 5 to 15 wt. % CaO (within such a range that total amount of MgO and CaO becomes 5 to 15 wt. %); 10 to 18 wt. % Na₂O; 0 to 5 wt. % K₂O (within such a range that total amount of Na₂O and K₂O becomes 10 to 20 wt. %); and 0 to 5 wt. % B₂O₃, and colorants including: 1.2 to 2.2 wt. % total iron oxide (T-Fe₂O₃) converted to Fe₂O₃; 0.001 to 0.03 wt. % CoO; 0 to 0.0008 % Se; and 0 to 0.2 % NiO.

Fig. 1.



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Description

BACKGROUND OF THE INVENTION AND RELATED ART STATEMENT

5 The present invention relates to an ultraviolet/infrared absorbent low transmittance glass. More particularly, it relates to an ultraviolet/infrared absorbent low transmittance glass which has an almost neutral color such as turquoise blue shade and deep green shade so that it is useful for windows of vehicles or buildings when used in combination with a glass having green shade with high visible light transmittance, and which has low or middle visible light transmittance, low total solar energy transmittance and low ultraviolet transmittance so that it is suitable for use as privacy glazing of
10 a vehicle.

Recently, a variety of glass with ultraviolet/infrared absorbability to be used as a vehicle windshield has been proposed with the view of preventing degradation of luxurious interior materials and reducing cooling load of the vehicle. In view of privacy protection, glass with relatively low visible light transmittance is preferably used for a rear window glass of a vehicle. Such kinds of glass include the followings.

15 For example, a dark gray colored infrared absorbent glass disclosed in Japanese Patent Publication No. 7-29813 consists of soda-lime-silica glass including colorants consisting of 1.00 to 1.7 weight percent Fe_2O_3 (total iron), at least 0.27 weight percent FeO , 0.002 to 0.005 weight percent Se , and 0.01 to 0.02 weight percent CoO . The glass exhibits luminous transmittance less than 32 percent and total solar infrared transmittance less than 15 percent at 3.9 mm thickness.

20 A dark gray colored glass disclosed in Japanese Unexamined Published Patent Application No. 8-157232 consists of soda-lime-silica glass including colorants consisting of 0.8 to 1.4 weight percent Fe_2O_3 (total iron), less than 0.21 weight percent FeO , 0.05 to 1.0 weight percent TiO_2 , 0.02 to 0.05 weight percent CoO , and 0.0005 to 0.015 weight percent Se .

A neutral gray colored glass disclosed in claim 25 of U.S. Patent No. 5,393,593 has a base glass composition comprising 66 to 75 weight percent SiO_2 , 10 to 20 weight percent Na_2O , 5 to 15 weight percent CaO , 0 to 5 weight percent MgO , 0 to 5 weight percent Al_2O_3 , and 0 to 5 weight percent K_2O , and colorants consisting of 1.00 to 2.2 weight percent Fe_2O_3 (total iron), at least 0.20 weight percent FeO , 0.0005 to 0.005 weight percent Se , and 0.010 to 0.030 weight percent CoO . The glass exhibits luminous transmittance less than 35 percent and total solar infrared transmittance less than 20 percent at 3.9 mm thickness.

30 A soda-lime-silica glass disclosed in Japanese Publication of PCT No. 8-506314 has a neutral color by including primary iron given by the following expression:

$$\text{FeO (weight percent)} \geq 0.007 + (\text{optical concentration} - 0.036) / 2.3$$

35 and 0.25 to 1.75 weight percent Fe_2O_3 and at least one selected from a group consisting of Se , Co_3O_4 , Nd_2O_3 , NiO , MnO , V_2O_5 , CeO_2 , TiO_2 , CuO , and SnO . The glass exhibits luminous transmittance more than 32 percent, ultraviolet transmittance less than 25 percent, direct solar energy transmittance at least 7 percent less than the luminous transmittance, and dominant wavelengths preferably less than 570 nanometers at 4 mm thickness. Certain embodiments thereof may be used as a privacy glazing.

40 In both the dark gray colored infrared absorbent glass disclosed in Japanese Patent Publication No. 7-29813 and the dark gray colored glass disclosed in Japanese Unexamined Published Patent Application No. 8-157232, a great quantity of Se is used for providing a desirable color. Such a great quantity of Se is unpreferable for the environment because Se has toxicity and is easy to vaporize. The above dark gray glass including 0.05 to 1.0 weight percent TiO_2 as an essential component is unpreferable because TiO_2 is expensive to increase the batch cost.

45 The neutral gray colored glass disclosed in U.S. Patent No. 5,393,593 is also unpreferable because of the great content of Se . The great content of FeO is preferable in view of heat ray absorptivity, FeO is on the contrary not preferable because it selectively absorbs infrared ray in a wavelength from 1,000 to 1,200 nanometers so that in case of the glass is produced by a normal melting furnace, it absorbs most efficient portion in flame luminance distribution so as to lower the temperature of bottom base of the furnace and then contribute to various defects.

50 Though the aforementioned glass with low visible light transmittance is superior in view of privacy protection, an occupant in a vehicle cabin sees outside quite unclearly through the glass. On the other hand, a glass with middle transmittance can be sufficient to some extent for both privacy protection and safety. These types of glass are chosen according to positions in a vehicle and circumferences.

The aforementioned glass includes selenium in high concentration to provide optical properties without essentially including nickel.

55 The glass of Japanese Publication PCT No. 8-506314, which can be used for privacy glazing, is provided with a neutral color by including colorants consisting all of nickel, selenium, and cobalt, just as described in this specification. However, it is necessary to add a great quantity of Se since the content of nickel is low.

OBJECT AND SUMMARY OF THE INVENTION

It is an object of the present invention to provide an ultraviolet/infrared absorbent low transmittance glass which has an almost neutral color such as turquoise blue and deep green and which has low or middle visible light transmittance, low ultraviolet transmittance, and low total solar energy transmittance.

The ultraviolet/infrared absorbent low transmittance glass of the present invention consists of a base glass, that is, the major constituents comprising:

- 65 to 80 wt. % SiO_2 ;
- 0 to 5 wt. % Al_2O_3 ;
- 0 to 10 wt. % MgO ;
- 5 to 15 wt. % CaO (within such a range that total amount of MgO and CaO becomes 5 to 15 wt. %);
- 10 to 18 wt. % Na_2O ;
- 0 to 5 wt. % K_2O (within such a range that total amount of Na_2O and K_2O becomes 10 to 20 wt. %); and
- 0 to 5 wt. % B_2O_3 , and a colorant including:
- 1.2 to 2.2 wt. % total iron oxide ($\text{T-Fe}_2\text{O}_3$) converted to Fe_2O_3 ;
- 0.001 to 0.03 wt. % CoO ;
- 0 to 0.0008 % Se ; and
- 0 to 0.2 % NiO .

BRIEF DESCRIPTION OF THE DRAWING

Fig. 1 is a diagram showing changes in color shade during reinforcing process in each example.

PREFERRED EMBODIMENTS

The description will be made as regard to an ultraviolet/infrared absorbent low transmittance glass composition. It should be noted that components will be represented with percentage by weight.

SiO_2 (silica) is a principal component for forming skeleton of glass. Less than 65% SiO_2 lowers the durability of the glass and more than 80% SiO_2 raises the melting temperature of the glass so high.

Al_2O_3 is a component for improving the durability of the glass. More than 5% Al_2O_3 raises the melting temperature of the glass so high. The preferable range of Al_2O_3 is between 0.1% and 2%.

MgO and CaO improve the durability of the glass and adjust a devitrification temperature and viscosity of the glass during molding. More than 10% MgO raises the devitrification temperature. Less than 5% or more than 15% CaO raises the devitrification temperature of the glass. The durability of the glass is lowered when the total amount of MgO and CaO is less than 5%, while the devitrification temperature is increased when the total exceeds 15%.

Na_2O and K_2O prompt the glass to melt. The efficiency of promotion of melting becomes poor when Na_2O is less than 10% or the total of Na_2O and K_2O is less than 10%, while the durability of the glass is lowered when Na_2O exceeds 18% or the total of Na_2O and K_2O exceeds 20%. K_2O is preferable not to exceed 5% because of its expensive cost.

B_2O_3 is a component for improving the durability of the glass, prompting to melt, and yet enhancing the ultraviolet absorption. B_2O_3 should be less than 5%, since the transmittance is reduced also at a visible range so that the color of the glass is easy to tint yellow and difficulties during molding are caused due to the vaporization of B_2O_3 when B_2O_3 exceeds 5%.

Iron oxide is present in the form of Fe_2O_3 and the form of FeO in the glass. Fe_2O_3 is a component for improving the ultraviolet absorptivity and FeO is a component for improving the heat ray absorptivity.

When the total amount of iron oxide calculated as Fe_2O_3 ($\text{T-Fe}_2\text{O}_3$) is less than 1.2%, the efficiency of ultraviolet and infrared absorptivity becomes small so as not to provide desired optical properties. On the other hand, when $\text{T-Fe}_2\text{O}_3$ exceeds 2.2%, there is an unpreferable possibility that the temperature around a crown of a glass melting furnace exceeds its refractory temperature due to absorption of the heat ray by the ferrous oxide. In addition, in case of successively producing glass by a glass melting furnace with $\text{T-Fe}_2\text{O}_3$ exceeding 2.2%, long time is required to alter a glass composition in the furnace. $\text{T-Fe}_2\text{O}_3$ content is preferably equal or more than 1.2% and less than 1.8% (more preferably between 1.25% and 1.35%) or between 1.8% and 2.2%.

Equal or more than 1.2% and less than 1.8% $\text{T-Fe}_2\text{O}_3$ can offer advantages of small load to the furnace because the glass composition in the successive glass melting furnace is altered to another composition within a short time. In this case, sufficient efficiency of ultraviolet absorptivity is sometimes not obtained only by iron. However, the efficiency of ultraviolet absorptivity can be improved by adding, for example, CeO_2 , TiO_2 within the ranges defined by the present invention.

Particularly 1.25% to 1.35% $\text{T-Fe}_2\text{O}_3$ can offer the aforementioned advantages and sufficient efficiency of ultraviolet

let absorptivity only by iron.

1.8% to 2.2% T-Fe₂O₃ makes the time longer to change the glass composition in the furnace, and the load to the furnace becomes greater. However, great efficiency of ultraviolet absorptivity can be obtained with low batch cost.

Fe₂O₃ has a function of particularly increasing the absorptivity in ultraviolet range when glass is reinforced by air blast cooling. This means that the glass of this invention can obtain enough efficiency of ultraviolet absorptivity without using expensive ultraviolet absorbent such as CeO₂ and TiO₂. When T- Fe₂O₃ is in the range mentioned above, the desired color shade of the glass can be obtained after discoloration due to the reinforcement process by air blast cooling.

The ratio (by weight) of FeO/T- Fe₂O₃ is preferably between 10% and 40%. When the ratio is less than 10%, enough heat ray absorptivity can not be obtained because of a small amount of FeO.

When the FeO/T- Fe₂O₃ ratio exceeds 40%, the visible light transmittance is reduced and the color is of a blue tint. In addition, nickel sulfide particles are sometimes present in the molten glass because of a great quantity of Fe²⁺ having reducing characteristics in the molten glass. The FeO/T- Fe₂O₃ ratio exceeding 40% causes streaks richly including silica and causes silica scum. The FeO/T- Fe₂O₃ ratio in a range between 10% and 40% brings a green shade which is an almost neutral color having high ultraviolet absorbability and high heat ray absorptivity. In this case, values converted to Fe₂O₃ are used for the content of FeO.

CoO is a component for forming an almost neutral color such as turquoise blue shade and deep green shade by cooperating with Se and/or NiO, and Fe₂O₃ and for controlling the visible light transmittance. Less than 0.001% CoO can not form a desired color shade and makes the visible light transmittance too high. More than 0.03% CoO makes the color of too blue tint and reduces the visible light transmittance.

Se contributes a pink color, so that it reduces the excitation purity with the aid of a complementary color of CoO. When NiO is included, Se may not be always included. More than 0.0008% Se reduces the visible light transmittance. When using Se, the content thereof is preferable in a range between 0.0001% and 0.0008%, particularly in a range between 0.0001 and 0.0004%. As mentioned, the desired color shade can be obtained with a significantly smaller amount of Se than the amount conventionally required or without Se. It should be noted that the remaining Se in the glass can be increased by setting the content of T- Fe₂O₃ and the FeO/T- Fe₂O₃ ratio in the aforementioned ranges.

NiO is a component for controlling the visible light transmittance and for reducing the excitation purity as like as CoO. It should be understood that NiO may not be always included when Se is included. When NiO is more than 0.2%, nickel sulfide stones are sometimes present in the product and the visible light transmittance is reduced. In addition, the obtained shade becomes too greenish. When using NiO, the content thereof is preferably in a range between equal to or more than 0.003% and less than 0.05% for middle visible light transmittance, or in a range between 0.05% and 0.2% for low visible light transmittance.

When the concentration of NiO in the glass is too high, there is a possibility that NiO coagulates to form a nickel sulfide stones. However, when NiO is in the range defined by this invention, the desired color shade can be obtained without producing the nickel sulfide stones.

It is known that the coordination number of NiO varies according to the rate of cooling glass so that the color of the glass varies. This is because the cooling varies the coordination number of oxide around Ni²⁺ from 6 into 4 and thus varies the optical absorption. The absorption of Ni²⁺ with 6 coordinating atoms exists around 430 nanometers so as to contribute yellow to the glass, while the absorption of Ni²⁺ with 4 coordinating atoms exists from 500 to 640 nanometers. Therefore, the excitation purity would be reduced to obtain the preferable shade by using Ni²⁺ with 4 coordinating atoms. Windshields of a passenger car are normally reinforced by air blast cooling for safety. The reinforcement process by air blast cooling also varies the coloring condition of NiO. In the present invention, the desired color shade can be obtained without adding Se by the discoloration due to the reinforcement process by air blast cooling.

CeO₂ is a component for improving the ultraviolet absorptivity and is present in the form of Ce³⁺ or in the form of Ce⁴⁺ in glass. Particularly, Ce³⁺ is effective in absorbing ultraviolet with less absorptivity in the visible range. In the present invention oxide of Ce³⁺ is also expressed in terms of CeO₂ and is included in the total amount of CeO₂.

TiO₂ is a component for improving the ultraviolet absorptivity particularly by interaction with FeO. TiO₂ can be added to improve the ultraviolet absorptivity within such a range as not to lose the almost neutral color such as turquoise blue shade and deep green shade, or to add a yellow tint in order to obtain the desired color shade. The use of expensive CeO₂, TiO₂ increases the cost so that it is not preferable to use more than 2% CeO₂, TiO₂.

One or more than two among MnO, V₂O₅, MoO₃, CuO, Cr₂O₃, and the like may be added as colorant and SnO₂ within a range from 0% to 1% in total may be added as a reducing agent in such a range as not to lose middle transmittance and the almost neutral color such as turquoise blue shade and deep green shade. To further securely prevent the formation of nickel sulfide stones, ZnO may be added in a range from 0% to 1%.

In the present invention, when measured using C.I.E. standard illuminant "A", the total solar energy transmittance (TG) is preferably lower than the visible light transmittance (YA). It is preferable that YA is in a range from 23 to 50% and TG is in a range from 7 to 35% when the measurement is made based on a glass thickness of 4 mm. More preferably, YA is in a range from 25 to 40% and TG is in a range from 20 to 35%. In case of using L* a* b* color system, the chro-

maticity, expressed by a^* and b^* , of the glass color are preferably in ranges of $-9 < a^* < -6$ and $-3 < b^* < 3$, respectively.

In the present invention, the glass preferably has YA between 10% and 25% and TG between 10% and 35% at a thickness from 3.1 to 5 mm. In this case, the chromaticity, expressed by a^* and b^* under $L^* a^* b^*$ color system, of the glass color is preferably in ranges of $-7 < a^* < -2$ and $-3 < b^* < 7$, respectively.

When measured using C.I.E. standard illuminant "C" over the wavelength range 380 to 770 nanometers, the glass of the present invention preferably has optical properties with dominant wavelengths in the range of 480 to 580 and excitation purity less than 11%.

The glass of the present invention preferably exhibits ultraviolet transmittance, defined by ISO 9050, less than 6%.

[Examples]

Hereinafter, the mode of carrying out the present invention will be described referring to some examples.

(Examples 1 through 35)

Glass raw material is prepared by adding at least one of a group consisting of ferric oxide, titanium oxide, cerium oxide, cobalt oxide, metallic selenium, and nickel oxide into a standard soda-lime-silica glass batch composition, also adding carbonaceous reducing agent (concretely, coke powder etc.) at a ratio of about 0.01 part by weight per 100 parts of the glass raw material therein, and mixing them. The glass raw material thus prepared is heated and melted in an electric furnace at 1500 °C for 4 hours. The molten glass is flowed onto a stainless plate and annealed to the room temperature as long as 16 hours to obtain a 6 mm thick glass plate. The glass plate is then polished in such a manner that the thickness reduces to 4 mm to become a sample. Each sample is measured in the visible light transmittance by the C.I.E. illuminant A(YA), the total solar energy transmittance (TG), the ultraviolet transmittance by ISO 9050(TUV), the dominant wavelength by the illuminant C(DW), and the excitation purity (Pe). Before measuring the optical properties, some of the samples are cooled by blasting air after re-heating in order to reinforce them.

Tables 1-4 show base glass compositions of the obtained sample, T-Fe₂O₃ concentration, FeO (converted to Fe₂O₃) / T-Fe₂O₃ rate, CoO concentration, Se concentration, NiO concentration, CeO₂ concentration, and TiO₂ concentration. The numerals in Tables are indicated as a percentage of the weight except that CoO concentration, Se concentration, and NiO concentration are expressed in ppm. The percentages of SiO₂ are indicated without figures after decimal points because the figures after decimal points are rounded off. Tables 1-4 also show the optical properties of the respective samples.

All of the samples of Examples 1 through 21 are characterized by visible light transmittance (YA) between 23% and 50%, total solar energy transmittance (TG) between 7% and 30%, and ultraviolet transmittance (TUV) less than 6%. These are also characterized by dominant wavelength (DW) measured using the illuminant C between 480 and 525 nanometers, and excitation purity less than 11%.

It is shown that Examples 2-6, 10, and 11 have the same or improved properties as the glass of Example 1, although these are free of NiO.

The examples 7-9, 12-16 have the same or improved properties as the glass of Example 1, although these are free of Se.

In Examples 2-9, the furnace's load during melting can be reduced because of small amounts of T-Fe₂O₃.

Examples 10-16 have improved ultraviolet absorptivity and improved infrared absorptivity because of large amounts of T-Fe₂O₃.

In order to enhance the ultraviolet absorptivity, relatively large amounts of TiO₂ are added in Examples 17-19, and CeO₂ is added in Example 20. These examples have large ultraviolet absorptivity without losing the visible light transmittance and total solar ultraviolet transmittance.

Example 21 is added CuO in order to reduce the total solar ultraviolet transmittance, and can exhibit high infrared absorptivity without losing the visible light transmittance and the ultraviolet absorptivity.

Examples 6, 7, and 14 have high visible light transmittance while having high infrared absorptivity and yet have high ultraviolet absorptivity so that these are superior in prevention of degradation of interior materials and reduction of cooling load of the vehicle. Therefore, these are suitable for the privacy protection type of visible light transmission glass, for example, rear windows of vehicles, windows of buildings, and the like.

The ultraviolet absorptivity and the infrared absorptivity of Examples 22-29 are slightly small. These samples contain smaller amount of T-Fe₂O₃ and the FeO/T-Fe₂O₃ ratio is set relatively low, so that they are superior in productivity.

Among these examples, Examples 23-29 have the same or improved properties as the glass of other examples, although these are free of Se.

Example 25 exhibits middle visible light transmittance and low infrared/ultraviolet transmittance.

Examples 26, 28, and 29 exhibit low visible light transmittance and low infrared/ultraviolet transmittance.

Examples 30-35 show results of measuring optical properties before and after the reinforcement process by air

blast cooling.

Among these examples, Examples 30-32 exhibit middle visible light transmittance and low infrared/ultraviolet transmittance and have an almost neutral color such as turquoise blue shade and deep green shade. The ultraviolet transmittance is improved about 1.5% by the reinforcement process.

Examples 33-35 exhibit low visible light transmittance and low infrared/ultraviolet transmittance. The ultraviolet transmittance is improved about 2% by the reinforcement process.

Each of Examples 30-35 has a preferable color shade which is an almost neutral color of gray tint through the reinforcement process. Fig. 1 shows a coordinate diagram of colors indicated by a^* , b^* for illustrating how points are moved by the reinforcement process in the aforementioned examples.

It would be apparent from the diagram that all of the points move toward the origin according to the reinforcement process so that the colors become closer to the desired color shade, i.e. almost neutral colors having reduced excitation purity.

Therefore, when the glass compositions of the examples mentioned above are used for windshields of vehicles and windows of building, good effects of preventing degradation of interior materials and of privacy protecting can be expected.

(Comparative Examples 1-4)

Table 5 shows glass components and optical properties of Comparative Examples which are made in the same manner as Examples 1-35 but the glass components are different.

All of Comparative Examples 1-4 have components out of the range of the present invention. Comparative Example 1 contains NiO as a colorant of which amount is out of the claimed range and Comparative Example 2 contains $T\text{-Fe}_2\text{O}_3$ of which amount is out of the claimed range. Comparative Example 3 has the same components as the example of Japanese Unexamined Published Patent Application No. H8-157232 as referred above and Comparative Example 4 has the same components as the example of U.S. Patent No. 5,393,593 as referred above. It should be noted that the optical properties of Comparative Example 3 are indicated in values converted based on a glass thickness of 5 mm and the optical properties of Comparative Example 4 are indicated in values converted based on a glass thickness of 3.9 mm.

It would be apparent from Table 5 that as compared with the examples of the present invention. Comparative Examples 1 and 2 do not have an almost neutral color such as turquoise blue shade and deep green shade and yet do not exhibit low or middle visible light transmittance, low infrared transmittance, and low ultraviolet transmittance. Comparative Examples 3 and 4 are not preferable in view of productivity and environment because large amounts of expensive and poisonous Se are used therein.

As detailed above, according to the present invention, an ultraviolet/infrared absorbent low transmittance glass, which exhibits low or middle visible light transmittance, low total solar energy transmittance, and low ultraviolet transmittance and which has an almost neutral color such as turquoise and deep green, can be provided.

The ultraviolet/infrared absorbent low transmittance glass having the almost neutral color such as turquoise and deep green can exhibit the effect of preventing degradation and discoloration of interior materials and the privacy protecting effect when the glass is used for a rear window glass of a vehicle, a window of a building, or the like.

[Table 1]

Example	1	2	3	4	5	6	7	8	9	11	12
SiO ₂	71	71	71	71	71	71	71	71	71	71	71
Al ₂ O ₃	1.6	1.6	1.6	1.6	1.6	1.6	1.6	1.6	1.6	1.6	1.6
MgO	3.6	3.6	3.6	3.6	3.6	3.6	3.6	3.6	3.6	3.6	3.6
CaO	7.7	7.7	7.7	8.4	7.7	7.7	7.7	7.7	7.7	7.7	7.7
Na ₂ O	13.7	13.7	13.7	13.1	13.7	13.7	13.7	13.7	13.7	13.7	13.7
K ₂ O	0.9	0.9	0.9	0.9	0.9	0.9	0.9	0.9	0.9	0.9	0.9
B ₂ O ₃	-	0.2	-	-	-	-	-	-	-	-	-
T-Fe ₂ O ₃	1.9	1.3	1.7	1.5	1.4	1.4	1.7	1.5	1.4	1.9	1.8
FeO/T-Fe ₂ O ₃	0.24	0.28	0.24	0.24	0.25	0.27	0.26	0.24	0.26	0.20	0.23
CeO ₂	-	-	-	-	-	-	-	-	-	-	-
TiO ₂	0.03	0.03	0.03	0.03	0.03	0.03	0.03	0.03	0.03	0.03	0.03
Se(ppm)	3	5	2	3	4	3	-	-	-	3	1
CoO(ppm)	100	95	113	107	101	92	94	96	98	125	119
NiO(ppm)	30	-	-	-	-	-	150	200	250	-	-
YA	30.8	38.2	34.1	35.1	37.0	36.2	35.4	38.2	37.5	33.1	32.9
TG	18.7	23.5	20.6	21.7	23.3	20.0	19.7	23.4	23.3	20.6	19.5
Tuv(ISO)	1.9	5.7	2.8	3.4	4.4	5.1	3.0	4.6	5.5	1.7	2.1
dominant wavelength Dw	519	492	494	496	495	498	499	498	497	499	495
excitation purity Pe	4.5	7.4	7.7	5.9	6.0	7.1	6.8	6.7	6.8	6.2	8.2

[Table 2]

Example	12	13	14	15	16	17	18	19	20	21
SiO ₂	70	70	70	70	70	71	71	69	70	71
Al ₂ O ₃	1.6	1.6	1.6	1.6	1.6	1.6	1.6	1.6	1.6	1.6
MgO	3.6	3.6	3.6	3.6	3.6	3.6	3.6	3.6	3.6	3.6
CaO	7.7	8.4	8.4	8.4	8.4	8.4	8.4	7.7	8.4	7.7
Na ₂ O	13.7	13.1	13.1	13.1	13.1	13.1	13.1	13.7	13.1	13.7
K ₂ O	0.9	0.9	0.9	0.9	0.9	0.9	0.9	0.9	0.9	0.9
B ₂ O ₃	-	-	-	-	0.5	-	-	-	-	-
T-Fe ₂ O ₃	2.2	2.2	1.9	1.8	1.3	1.3	1.5	1.4	1.3	1.3
FeO/T-Fe ₂ O ₃	0.34	0.30	0.24	0.22	0.31	0.28	0.29	0.30	0.40	0.23
CeO ₂	-	-	-	-	-	-	-	-	1	-
TiO ₂	0.03	0.03	0.03	0.03	0.5	0.5	0.5	1.6	0.03	0.03
Se(ppm)	-	-	-	-	-	5	-	2	-	-
CoO(ppm)	98	100	90	92	130	95	75	115	75	110
NiO(ppm)	300	300	50	100	300	-	500	-	500	-
others										CuO 0.22
YA	23.0	23.2	35.6	37.7	31.9	38.0	30.4	28.0	30.4	43.9
TG	10.2	11.1	19.3	21.8	20.5	23.2	15.4	16.1	15.4	28.1
Tuv(ISO)	0.9	1.0	1.9	2.6	4.0	3.7	3.9	1.6	3.9	3.8
dominant wavelength Dw	511	517	499	500	497	516	520	525	514	492
excitation purity Pe	7.3	6.6	7.4	6.6	7.5	4.1	8.3	9.8	8.3	9.6

[Table 3]

Example	22	23	24	25	26	27	28	29
SiO ₂	71	71	71	70	71	71	71	71
Al ₂ O ₃	1.6	1.6	1.6	1.6	1.6	1.6	1.6	1.6
MgO	3.6	3.6	3.6	3.6	3.6	3.6	3.6	3.6
CaO	7.7	7.7	7.7	7.7	7.7	7.7	7.7	7.7
Na ₂ O	13.7	13.7	13.7	13.7	13.7	13.7	13.7	13.7
K ₂ O	0.9	0.9	0.9	0.9	0.9	0.9	0.9	0.9
B ₂ O ₃	-	-	-	-	-	-	-	-
T-Fe ₂ O ₃	1.2	1.4	1.7	1.3	1.3	1.4	1.3	1.3
FeO/T-Fe ₂ O ₃	0.24	0.21	0.19	0.19	0.16	0.17	0.17	0.17
CeO ₂	-	-	-	0.4	-	-	-	-
TiO ₂	0.03	0.03	0.03	0.03	0.50	0.03	0.03	0.03
Se(ppm)	3	-	-	-	-	-	-	-
CoO(ppm)	110	130	100	120	300	80	180	200
NiO(ppm)	-	390	250	300	600	1900	1050	1000
YA	37.9	35.4	37.4	38.3	18.7	17.7	20.2	19.9
TG	21.2	28.4	24.6	29.8	25.2	21.8	21.1	23.2
Tuv(ISO)	5.6	8.3	2.6	3.6	4.9	8.1	6.4	6.8
dominant wavelength Dw	496	496	510	499	485	565	555	543
excitation purity Pe	5.9	6	4.7	5.3	16.1	22.9	9.7	5.9

[Table 4]

Example	30		31		32		33		34		35	
	before	after	before	after	before	after	before	after	before	after	before	after
SiO ₂	71		71		71		71		71		71	
Al ₂ O ₃	1.6		1.6		1.6		1.6		1.6		1.6	
MgO	3.6		3.6		3.6		3.6		3.6		3.6	
CaO	7.7		7.7		7.7		7.7		7.7		7.7	
Na ₂ O	13.7		13.7		13.7		13.7		13.7		13.7	
K ₂ O	0.9		0.9		0.9		0.9		0.9		0.9	
B ₂ O ₃	—		—		—		—		—		—	
NiO (ppm)	330		330		330		1000		1000		900	
Se (ppm)	—		—		—		—		—		—	
CoO (ppm)	120		125		120		190		180		190	
TiO ₂	0.03		0.03		0.03		0.03		0.03		0.03	
t-Fe ₂ O ₃	1.3		1.3		1.3		1.3		1.3		1.3	
FeO/T-Fe ₂ O ₃	0.17		0.17		0.17		0.16		0.18		0.19	
YA TG	before	after	before	after	before	after	before	after	before	after	before	after
	38.8	37.4	37.2	35.9	37.8	36.3	20.2	17.8	20.2	17.7	20.4	18.1
	31.4	31.2	30.5	30.5	30.3	29.9	21.7	21.2	20.3	19.7	19.8	19.3
L* a* b*	69.50	68.38	68.37	67.25	68.79	67.59	52.54	49.63	52.63	49.54	53.01	50.27
	-9.47	-8.31	-9.45	-8.32	-9.48	-8.36	-9.10	-6.27	-9.39	-6.25	-9.64	-6.76
	0.78	0.59	0.14	0.01	-0.31	-0.34	6.72	3.77	6.19	2.96	3.39	0.69
dominant wavelength excitation purity P _e	498.9	498.5	496.9	496.5	495.8	495.5	550.3	542.4	545.9	529.1	514.9	499.5
	4.89	4.43	5.56	5.08	5.99	5.42	8.16	4.46	7.22	3.28	4.10	4.36
Tuv (ISO)	6.73	5.06	6.35	4.73	6.99	5.14	5.86	4.18	6.78	4.84	6.56	4.74

[Table 5]

Comparative Example	1	2	3	4
SiO ₂	71	71	72	66
Al ₂ O ₃	1.6	1.6	1.7	3.1
MgO	3.6	3.6	4.1	3.1
CaO	8.4	8.4	7.8	7.9
Na ₂ O	13.1	13.1	13.5	17.8
K ₂ O	0.9	0.9	0.5	-
B ₂ O ₃	-	-	-	-
T-Fe ₂ O ₃	1.4	1.0	1.4	1.5
FeO/T-Fe ₂ O ₂	0.61	0.21	0.17	0.18
CeO ₂	-	-	-	-
TiO ₂	0.03	0.03	0.1	-
Se(ppm)	-	-	43	40
CoO(ppm)	135	150	235	213
NiO	2200	300	-	-
thickness(mm)	4	4	5	3.9
YA	12.6	40.0	17.1	15.6
TG	7.2	33.1	16.6	16.7
Tuv(ISO)	8.8	21.7	2.5	4.7
dominant wavelength Dw	587	483	530	576
excitation purity Pe	18.0	20.5	3.9	8.9

Claims

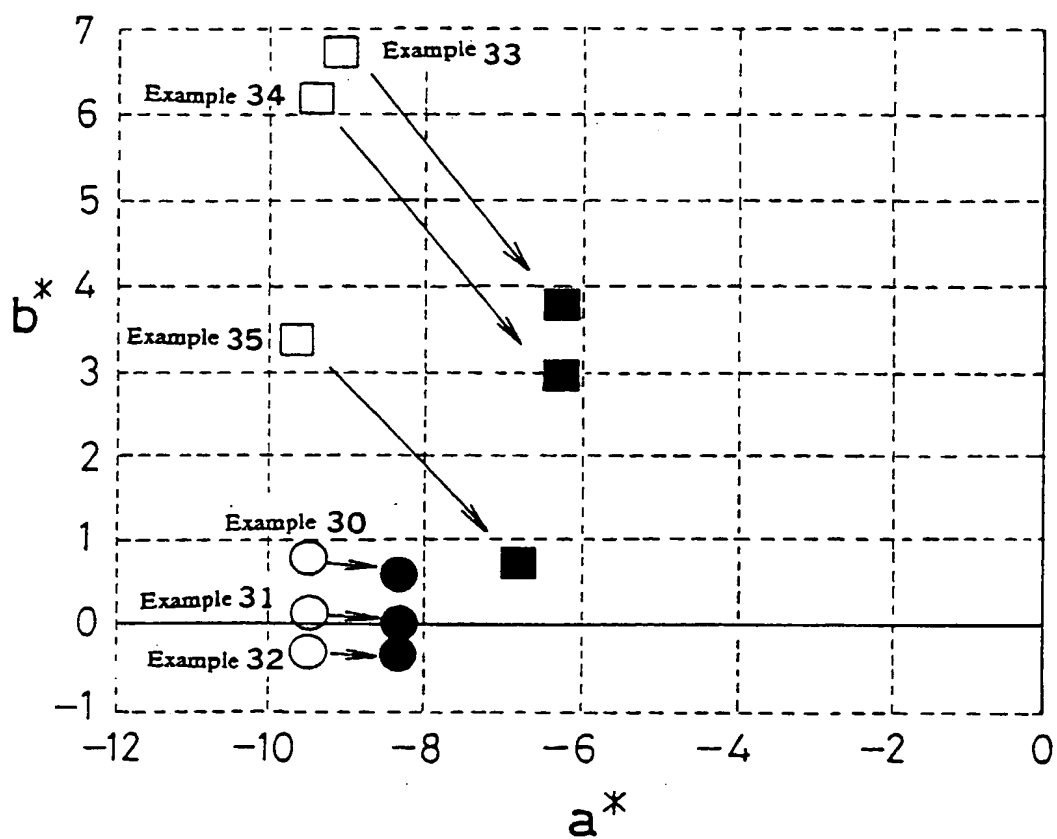
1. An ultraviolet/infrared absorbent low transmittance glass consisting of base glass comprising:

65 to 80 wt. % SiO₂;
 0 to 5 wt. % Al₂O₃;
 0 to 10 wt. % MgO;
 5 to 15 wt. % CaO (within such a range that the total amount of MgO and CaO becomes 5 to 15 wt. %);
 10 to 18 wt. % Na₂O;
 0 to 5 wt. % K₂O (within such a range that the total amount of Na₂O and K₂O becomes 10 to 20 wt. %); and
 0 to 5 wt. % B₂O₃, and colorants comprising:
 1.2 to 2.2 wt. % total iron oxide (T-Fe₂O₃) converted to Fe₂O₃;
 0.001 to 0.03 wt. % CoO;
 0 to 0.0008 % Se; and
 0 to 0.2 % NiO.

2. An ultraviolet/infrared absorbent low transmittance glass as claimed in claim 1, wherein T-Fe₂O₃ is equal or greater than 1.2 wt. % and less than 1.8 wt. %.
3. An ultraviolet/infrared absorbent low transmittance glass as claimed in claim 1, wherein T-Fe₂O₃ is between 1.8 wt. % and 2.2 wt. %.
4. An ultraviolet/infrared absorbent low transmittance glass as claimed in any of claims 1 to 3, wherein CoO is

between 0.001 wt. % to 0.018 wt. %.

5. An ultraviolet/infrared absorbent low transmittance glass as claimed in any of claims 1 to 4, wherein Se is between 0.0001 wt. % to 0.0008 wt. %.
6. An ultraviolet/infrared absorbent low transmittance glass as claimed in claim 5, wherein Se is between 0.0001 wt. % to 0.0004 wt. %.
7. An ultraviolet/infrared absorbent low transmittance glass as claimed in any of claims 1 to 4, wherein Se is not substantially comprised.
8. An ultraviolet/infrared absorbent low transmittance glass as claimed in any of claims 1 to 7, wherein NiO is between 0.003 wt. % to 0.2 wt. %.
9. An ultraviolet/infrared absorbent low transmittance glass as claimed in claim 8, wherein NiO is equal or greater than 0.003 wt. % and less than 0.05 wt. %.
10. An ultraviolet/infrared absorbent low transmittance glass as claimed in claim 8, wherein NiO is between 0.05 wt. % to 0.2 wt. %.
11. An ultraviolet/infrared absorbent low transmittance glass as claimed in any of claims 1 to 6, wherein NiO is not substantially comprised.
12. An ultraviolet/infrared absorbent low transmittance glass as claimed in any of claims 1 to 11, wherein said colorant further comprises CeO₂ of no greater than 2.0 wt. % and/ or TiO₂ of no greater than 2.0 wt. %.
13. An ultraviolet/infrared absorbent low transmittance glass as claimed in any of claims 1 to 12, wherein the amount of FeO calculated as Fe₂O₃ is in the range from 15 to 40% of T-Fe₂O₃.
14. An ultraviolet/infrared absorbent low transmittance glass as claimed in any of claims 1 to 13, wherein a total solar energy transmittance (TG) of said glass is smaller than a visible light transmittance (YA) by the C.I.E. illuminant A, and YA is in a range from 23% to 50% and TG is in a range from 7% to 35% at 4.0 mm thickness.
15. An ultraviolet/infrared absorbent low transmittance glass as claimed in claim 14, wherein YA is between 25% and 40% and TG is between 20% and 35%.
16. An ultraviolet/infrared absorbent low transmittance glass as claimed in claim 14 or 15, wherein the chromaticity, expressed by a* and b* under L* a* b* color system, of the glass color are in ranges of -9<a*<-6 and -3<b*<3, respectively.
17. An ultraviolet/infrared absorbent low transmittance glass as claimed in any of claims 1 to 16, said glass has YA between 10% and 25% and TG between 10% and 35% at a thickness between 3.1 mm and 5 mm.
18. An ultraviolet/infrared absorbent low transmittance glass as claimed in claim 17, wherein the chromaticity, expressed by a* and b* under L* a* b* color system, of the glass color are in ranges of -7<a*<-2 and -3<b*<7, respectively.
19. An ultraviolet/infrared absorbent low transmittance glass as claimed in any of claims 1 to 18, wherein said glass has an ultraviolet transmittance, defined by ISO, of no greater than 6%.
20. An ultraviolet/infrared absorbent low transmittance glass as claimed in any of claims 1 to 19, wherein the dominant wavelength of the glass is in the range of 480 to 525 nanometers when the measurement is made based on a glass thickness of 4 mm using illuminant C.
21. An ultraviolet/infrared absorbent low transmittance glass as claimed in any of claims 1 to 20, wherein the excitation purity of the glass is less than 11% when the measurement is made based on a glass thickness of 4 mm using illuminant C.

Fig. 1.



DOCUMENTS CONSIDERED TO BE RELEVANT							
Category	Citation of document with indication, where appropriate, of relevant passages	Relevant to claim	CLASSIFICATION OF THE APPLICATION (Int.Cl.6)				
X	US 5 545 596 A (ALVAREZ CASARIEGO PEDRO ET AL) 13 August 1996 * examples 2, 3, 7, 10, 12, 14 * * column 3 *	1-7, 11-24	C03C3/087 C03C4/02 C03C4/08				
X	US 5 308 805 A (BAKER RODNEY G ET AL) 3 May 1994 * column 2, line 11 - column 5, line 37 *	1-24					
P,X	WO 97 17303 A (ASAHI GLASS CO LTD ;SASAGE MIZUKI (JP); KIJIMA TAKASHI (JP); TANII) 15 May 1997 * abstract: examples 15-20,37-44 *	1,2,4-7, 11-24					
P,X	JP 08 245 238 A (NIPPON SHEET GLASS CO LTD) 24 September 1996 * abstract: examples 1-18 *	1-4,7, 11-24					
P,X	DE 196 36 303 A (GLAVERBEL) 13 March 1997 * page 3, line 38 - page 5, line 22; examples 7,16,17 *	1,2,4-6, 11-24	TECHNICAL FIELDS SEARCHED (Int.Cl.6)				
A	EP 0 452 207 A (SAINT GOBAIN VITRAGE) 16 October 1991 * the whole document *	1-24	C03C				
D,A	US 5 393 593 A (GULOTTA JOSEPH A ET AL) 28 February 1995 * column 3, line 57 - column 6, line 42; tables I,II *	1-24					
The present search report has been drawn up for all claims							
Place of search THE HAGUE		Date of completion of the search 19 November 1997	Examiner Van Bomme1, L				
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